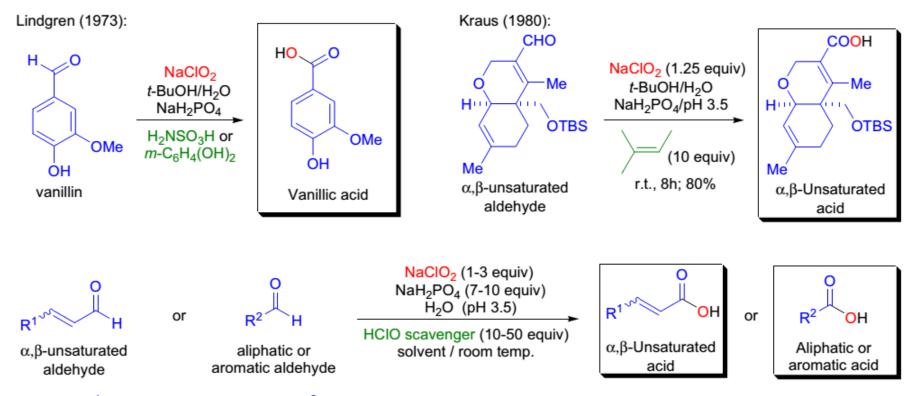
PINNICK OXIDATION



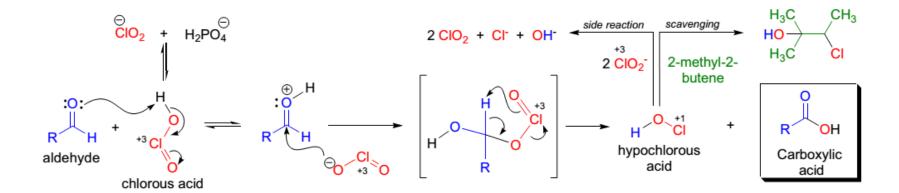
 R^1 = H, alkyl, aryl, alkenyl, allyl; R^2 = alkyl, aryl, allyl, homoallyl; <u>scavenger</u> = 2-methyl-2-butene, H₂O₂, H₂NSO₃H, m-C₆H₄(OH)₂, DMSO; <u>solvent</u> = t-BuOH, t-BuOH/THF

Features

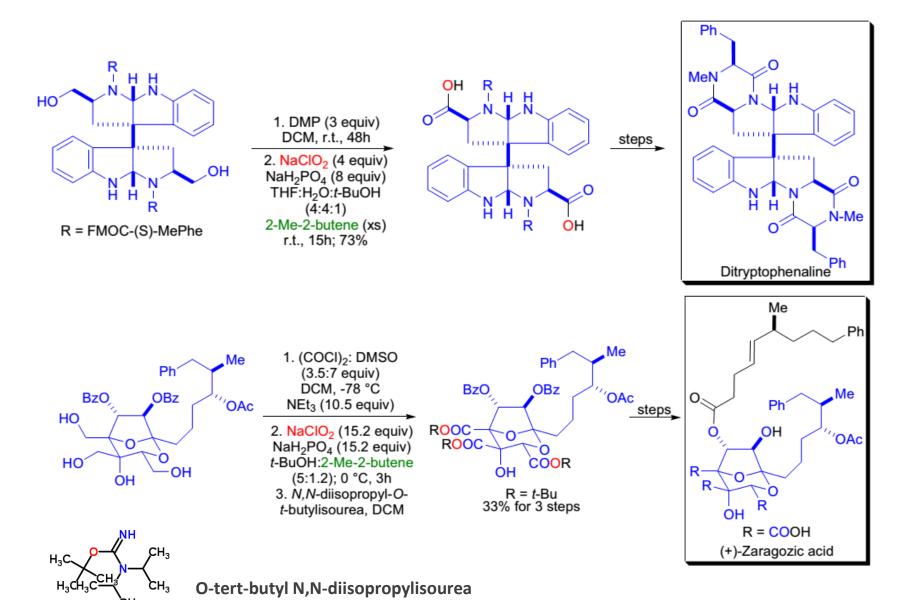
 R^1 = H, alkyl, aryl, alkenyl, allyl; R^2 = alkyl, aryl, allyl, homoallyl; <u>scavenger</u> = 2-methyl-2-butene, H₂O₂, H₂NSO₃H, m-C₆H₄(OH)₂, DMSO; <u>solvent</u> = t-BuOH, t-BuOH/THF

- ✓ In a typical procedure, the aldehyde is dissolved in tert-butanol (often in combination with another solvent such as THF) along with the large excess of the scavenger followed by the dropwise addition of the aqueous solution of sodium dihydrogen phosphate buffer (NaH₂PO₄) and NaClO₂ at room temperature.
- ✓ To ensure a constant pH value, the use of several equivalents of NaH2PO₄ is recommended.
- ✓ With certain substrates the purity of the reagents is crucial, and the oxidation sometimes stops after a few percent of conversion.
- ✓ Due to the sensitivity/instability of the NaClO₂ in acidic medium in the presence of transition metal complexes the use of a steel needle for the addition of the oxidant should be avoided (use a Pasteur pipette instead).
- \checkmark Stereocenters at the α -position of aldehydes are unaffected.
- ✓ Functional group tolerance is excellent, and hydroxyl groups do not need to be protected.

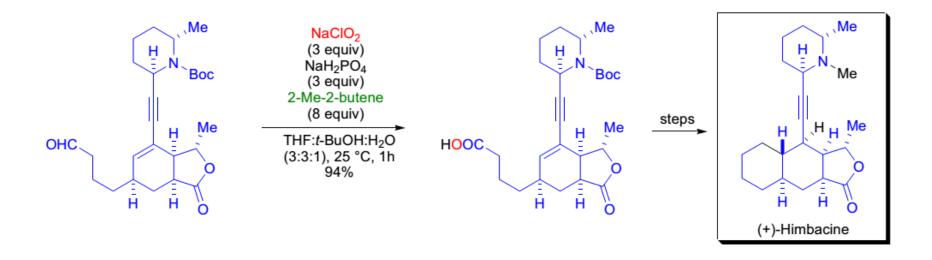
Mechanism



Synthetic Applications



Synthetic Applications



Thanks